

# Small-Scale Lab: Names and Formulas of Ionic Compounds

Name: \_\_\_\_\_

Period: \_\_\_\_\_

## Safety:

Wear safety goggles and follow the standard safety procedures.

## Purpose:

To observe the formation of compounds, and to write their names and formula:

## Materials:

- pencil
- paper
- reaction surface
- chemicals shown in figure A

## Procedure:

Place one drop of each chemical on the reaction surface in the squares designated by figure A. On your paper, record your observations in figure A for each solution.

	AgNO <sub>3</sub> (Ag <sup>+</sup> )	Pb(NO <sub>3</sub> ) <sub>2</sub> (Pb <sup>2+</sup> )	CaCl <sub>2</sub> (Ca <sup>2+</sup> )
Na <sub>2</sub> CO <sub>3</sub> (CO <sub>3</sub> <sup>2-</sup> )	a	e	i
Na <sub>3</sub> PO <sub>4</sub> (PO <sub>4</sub> <sup>3-</sup> )	b	f	j
NaOH (OH <sup>-</sup> )	c	g	k
Na <sub>2</sub> SO <sub>4</sub> (SO <sub>4</sub> <sup>2-</sup> )	d	h	l

Figure A

## Analysis:

Using the experimental data, record the answers to the following questions.

Describe each precipitate that forms as milky, grainy, cloudy, or gelatinous. Which mixture(s) did not form a precipitate? Write the formulas and names of the chemical compounds produced in the mixings.

Description	Formula	Name
a. _____	_____	_____
b. _____	_____	_____
c. _____	_____	_____
d. _____	_____	_____
e. _____	_____	_____
f. _____	_____	_____
g. _____	_____	_____
h. _____	_____	_____
i. _____	_____	_____
j. _____	_____	_____
k. _____	_____	_____
l. _____	_____	_____